

# **Chapter 9 Review**

# **Stoichiometry Answers**

## **Section 1**

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## Chapter 9 Review Stoichiometry

### Answers CHAPTER 9 REVIEW

### Stoichiometry MIXED REVIEW

### SHORT ANSWER Answer the

following questions in the space provided. 1. Given the following

equation:  $C_3H_4(g) + xO_2(g) \rightarrow$

$3CO_2(g) + 2H_2O(g)$  4 a. What is

the value of the coefficient x in this

equation? 40.07 g/mol b. What is

the molar mass of  $C_3H_4$ ? 2 mol  $O_2$ :

1 mol  $H_2O$  c. What is the mole ratio of  $O_2$  to  $H_2O$

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number is: Q. Using the pictured

equation, how many grams of zinc

chloride are produced from 7.89

moles of zinc? Chapter 9

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Quizlet Chapter 9: Standard Review

Worksheet 1. Answers will vary. An

example is included below:  $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$  This

describes the decomposition

reaction of hydrogen peroxide.

Microscopic: Two molecules of

hydrogen peroxide (in aqueous

solution) decompose to produce

two molecules of liquid water and

one molecule of oxygen

gas. Chapter 9: Standard Review

Worksheet Modern Chemistry 77

Stoichiometry CHAPTER 9 REVIEW

Stoichiometry SECTION 3

PROBLEMS Write the answer on the

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line to the left. Show all your work in the space provided. 1. \_\_\_\_\_ The actual yield of a reaction is 22 g and the theoretical yield is 25 g.

Calculate the percentage yield. 2.

6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub>

### CHAPTER 9 REVIEW

### Stoichiometry Chapter 9 -

Stoichiometry Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of

reactants/products involved in a...

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### Stoichiometry. MIXED REVIEW.

### SHORT ANSWER Answer the

following questions in the space provided. 1. Given the following

equation:  $C_3H_4(g) + x \cdot O_2(g) \rightarrow$

$3CO_2(g) + 2H_2O(g)$  a. What is the

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### Stoichiometry Answers Section 1

value of the coefficient  $x$ . in this equation? b. What is the molar mass of  $C_3H_4$ ? c. How many moles are in an 8.0 g sample of  $C_3H_4$ ? 2.

a. What is meant by ideal conditions CHAPTER 9 REVIEW -

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answers to Chapter 3 -

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Theoretically, how many moles of  $\text{NH}_3$  will be produced? PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of  $\text{N}_2$  are mixed with 12.0 mol of  $\text{H}_2$  according to the ... Date. FCHAPJ REV[EW. Chapter 9 Mixed Review Stoichiometry Answers Recognizing

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Stoichiometry. SECTION 1. SHORT ANSWER Answer the following

questions in the space provided. 1.

\_\_\_\_\_ The coefficients in a chemical equation represent the. (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

(c) number of atoms of each element in each compound in a reaction. CHAPTER 9 REVIEW -

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Resources - Home Reaction

stoichiometry, the subject of this

chapter, is based on chemical

equations and the law of

conservation of mass. All reaction

stoichiometry ... 290 Chapter 9 DO

NOT EDIT--Changes must be made

through "File info" ... The number of

significant figures in the answer



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